

GENERAL SOLUBILITY RULES

ION	RULE
S^{2-}	All sulfides are insoluble; the only exceptions are K^+ , Na^+ & NH_4^+ Ca^{2+} , Mg^{2+} , Ba^{2+} & Sr^{2+} (alkaline earths).
NO_3^-	All nitrates are soluble.
OH^-	All hydroxides are insoluble; exceptions are K^+ & Na^+ . Ca^{2+} & Ba^{2+} hydroxides are slightly soluble.
SO_4^{2-}	All sulfates are soluble, aside from Pb^{2+} , Sr^{2+} , Ba^{2+} & Ca^{2+} .
PO_4^{3-} & CO_3^{2-}	All phosphates and carbonates are insoluble, except K^+ , NH_4^+ & Na^+ .
K^+ , Na^+ & NH_4^+	All salts of potassium ion, sodium ion and ammonium ion are soluble; the only exceptions are a few that are uncommon.
I^- , Br^- & Cl^-	All iodides, bromides and chlorides are soluble; exceptions are Pb^+ , Hg_2^{2+} & Ag^+ .
$C_2H_3O_2^-$	All acetates are soluble, with the exception of $AgC_2H_3O_2$, which is only slightly soluble.